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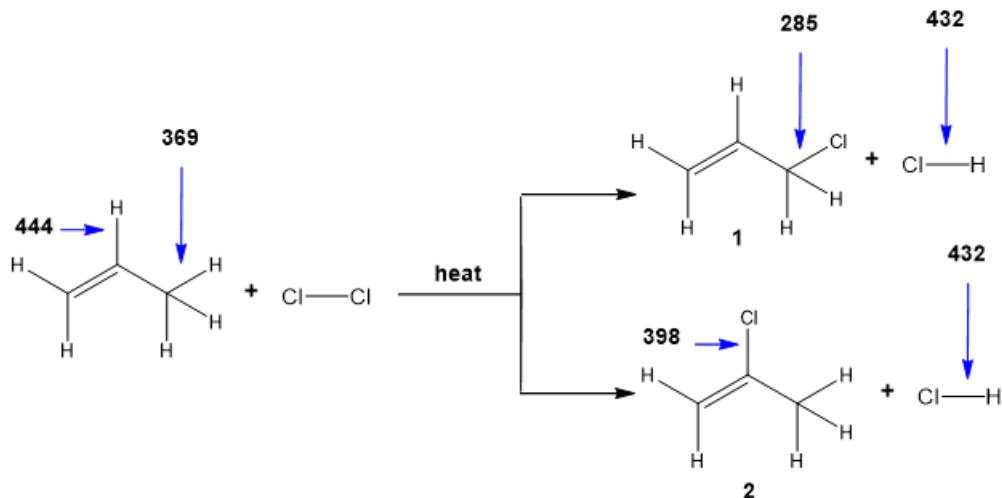
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The chlorination of propene can give the products shown below. The bond dissociation energies (kJ/mol) of the bonds are indicated using arrows.



- Calculate the overall ΔH^\ddagger for the formation of 3-chloropropene (1) and 2-chloropropene (2).
- Calculate the ΔH^\ddagger for the rate determining step in the chlorination of 3-chloropropene (1) and 2-chloropropene (2).
- Identify the most thermodynamically stable product and the kinetic product of this reaction

VIDEO SOLUTION

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How to Reach Us

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