



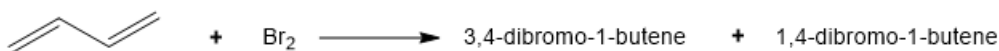
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Reaction of 1, 3-butadiene with one equivalent of bromine gives a mixture of two products in different ratios at -15°C and 60°C as shown. Select the correct statement regarding this reaction.



-15°C 60% A, 40% B

60°C 10% A, 90% B

- ☐ (A) Product A is 3, 4-dibromo-1-butene, which is more stable than 1, 4-dibromo-2-butene (B) at -15°C but is less stable than B at 60°C .
- ☐ (B) Product A is 1, 4-dibromo-2-butene, which is more stable than 3, 4-dibromo-1-butene (B) at -15°C but is less stable than B at 60°C .
- ☐ (C) Product A is 3, 4-dibromo-1-butene, which has a lower activation energy for formation than 1, 4-dibromo-2-butene (B) at -15°C but has a higher activation energy for formation than B at 60°C .
- ☐ (D) Product A is 1, 4-dibromo-2-butene, which has a lower activation energy for formation than 3, 4-dibromo-1-butene (B) at -15°C but has a higher activation energy for formation than B at 60°C .
- ☐ (E) Product A is 3, 4-dibromo-1-butene, which has a lower activation energy for formation than 1, 4-dibromo-2-butene (B) but is less stable than B.
- ☐ (F) Product A is 1, 4-dibromo-2-butene, which has a lower activation energy for formation than 3, 4-dibromo-1-butene (B) but is less stable than B.

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