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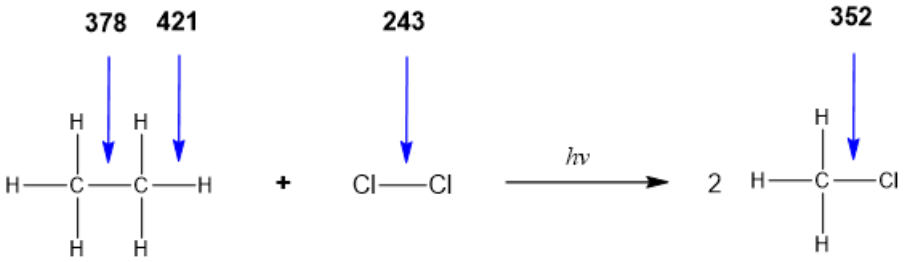
Calculate the ΔHËš for the following reaction from the bond dissociation energies (kJ/mol) shown below.



- **(A)** ΔHËš = +269 kJ/mol.
- **(B)** ΔHËš = -83 kJ/mol.
- **(C)** ΔHËš = -352 kJ/mol
- **(D)** ΔHËš = -41kJ/mol.
- **(E)** ΔHËš = +41 kJ/mol.

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