



ORGOREVIEW

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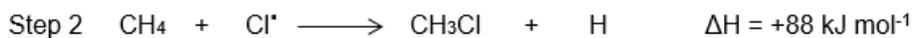
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In class we discussed the chlorination of methane. Here is an alternate mechanism.

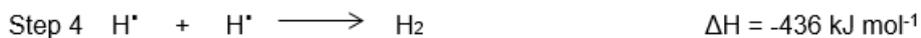
Chain initiation



Chain Propagation



Chain Termination



Bond Energies

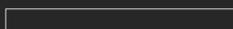
Cl-Cl	243 kJ mol ⁻¹
H-H	436 kJ mol ⁻¹
H ₃ C-Cl	352 kJ mol ⁻¹
H ₃ CH	440 kJ mol ⁻¹

This mechanism is thought to be *incorrect*. What is the best reason for why it is incorrect?

- (A) Step 1, the chain initiation step, is highly endothermic.

- (B) Step 2, the first chain propagation step, is highly endothermic.
- (C) Step 3, the second chain propagation step, is highly exothermic.
- (D) Steps 4, 5, and 6, the chain termination steps are all highly exothermic.
- (E) The overall reaction is exothermic.

VIDEO SOLUTION



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