



ORGOREVIEW

Question Vault

Welcome to orgoreview question vault. We have 5000 more problems for you to solve.

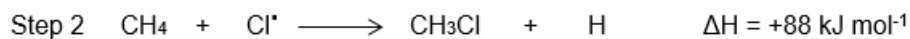
Share

In class we discussed the chlorination of methane. Here is an alternate mechanism.

Chain initiation



Chain Propagation



Chain Termination



Bond Energies

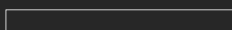
Cl-Cl	243 kJ mol ⁻¹
H-H	436 kJ mol ⁻¹
H ₃ C-Cl	352 kJ mol ⁻¹
H ₃ CH	440 kJ mol ⁻¹

This mechanism is thought to be *incorrect*. What is the best reason for why it is incorrect?

- ☐ (A) Step 1, the chain initiation step, is highly endothermic.

- ☐ (B) Step 2, the first chain propagation step, is highly endothermic.
- ☐ (C) Step 3, the second chain propagation step, is highly exothermic.
- ☐ (D) Steps 4, 5, and 6, the chain termination steps are all highly exothermic.
- ☐ (E) The overall reaction is exothermic.

VIDEO SOLUTION



How to Reach Us

Todd's Test Prep
2255 Glades Road
Suite 324A
Boca Raton, Florida 33431
E-mail: help@orgoreview.com

Links

[Contact](#)
[Terms](#)
[Privacy Policy](#)

follow us



Copyright 2025. All Rights Reserved.