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Each of the following reactions is in water. Which one of the following reactions is spontaneous ($\Delta G^\circ < 0$) as drawn?

(Hint: think about the pKa's of the conjugate acids of the bases in these reactions).

(A) $\text{CH}_3\text{I} + \text{NH}_3 \rightleftharpoons \text{CH}_3\text{NH}_3^+ + \text{I}^-$

(B) $\text{CH}_3\text{SH} + \text{NaCl} \rightleftharpoons \text{CH}_3\text{Cl} + \text{NaSH}$

(C) $\text{CH}_3\text{CH}_2\text{OH} + \text{NaBr} \rightleftharpoons \text{CH}_3\text{CH}_2\text{Br} + \text{NaOH}$

(D) $\text{CH}_3\text{CH}_2\text{NH}_2 + \text{NaI} \rightleftharpoons \text{CH}_3\text{CH}_2\text{I} + \text{NaNH}_2$

(E) None of the above

(A) (B) (C) (D) (E)

How to Reach Us

Todd's Test Prep

2255 Glades Road
Suite 324A
Boca Raton, Florida 33431
E-mail: help@orgoreview.com

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