



ORGOREVIEW

Question Vault

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Each of the following reactions is in water. Which one of the following reactions is spontaneous ($\Delta G^\circ < 0$) as drawn?

(Hint: think about the pKa's of the conjugate acids of the bases in these reactions).

- (A) $\text{CH}_3\text{I} + \text{NH}_3 \rightleftharpoons \text{CH}_3\text{NH}_3^+ + \text{I}^-$
- (B) $\text{CH}_3\text{SH} + \text{NaCl} \rightleftharpoons \text{CH}_3\text{Cl} + \text{NaSH}$
- (C) $\text{CH}_3\text{CH}_2\text{OH} + \text{NaBr} \rightleftharpoons \text{CH}_3\text{CH}_2\text{Br} + \text{NaOH}$
- (D) $\text{CH}_3\text{CH}_2\text{NH}_2 + \text{NaI} \rightleftharpoons \text{CH}_3\text{CH}_2\text{I} + \text{NaNH}_2$
- (E) None of the above

☐ (A) ☐ (B) ☐ (C) ☐ (D) ☐ (E)

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