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Using the following pKa values for lysine shown below give the structure of the product when lysine, as the anion, reacts with one equivalent acid to give the neutral form of lysine.

$$pK_{a} = 9.0 \xrightarrow{\text{NH}_{3}} \text{OH}$$

$$pK_{a} = 10.5 \xrightarrow{\text{OH}} \text{pK}_{a} = 2.2$$

$$\text{lysine}$$

VIDEO SOLUTION

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