



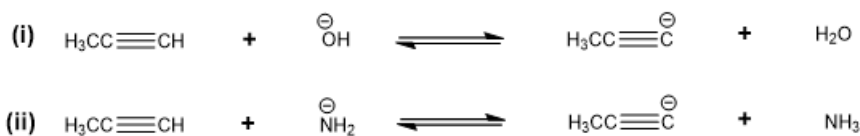
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Choose the incorrect statement about the following acid/base reactions involving propyne and its anion propynide.



	pK _a
H ₂ O	16
H ₃ CC≡CH	25
NH ₃	34

- (A) The equilibrium in equation (i) lies to the left. (B) The equilibrium in equation (ii) lies to the right.
- (C) You can prepare propynide, CH₃C≡C⁻, salts in NH₃.
- (D) Propyne, CH₃C≡CH, is a stronger acid than ammonia, NH₃.
- (E) In reaction (ii) ammonia, NH₃, acts as a base.

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