



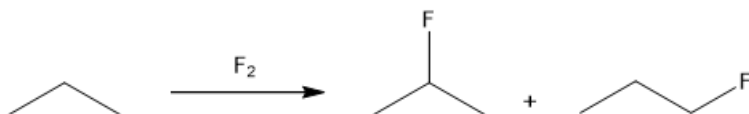
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The fluorination of propane is very highly exothermic.



Which one of the following statements is *not* true?

- ☐ (A) One would expect a near random substitution of the various hydrogen atoms in propane.
- ☐ (B) Only a small amount of 1-fluoropropane will be formed.
- ☐ (C) According to Hammond's postulate, the structures of the transition states would be similar to the structure of the reactant.
- ☐ (D) Heat will be released when fluorine reacts with propane.
- ☐ (E) One would expect the regioselectivity of fluorination to be poor.

VIDEO SOLUTION



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