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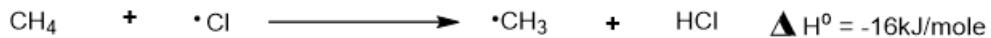
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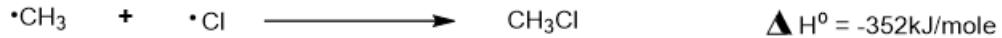
Choose the **incorrect** statement about the generally accepted free radical reaction mechanism for the chlorination of methane.



a) The change in enthalpy of the following propagation step is -16 kJ/mol.



b) The change in enthalpy for the following termination step is -352 kJ/mol.



c) Chlorine atoms are produced in the initiation step.

d) Methyl radicals are produced in one of the propagation steps.

e) Only a small fraction of the chloromethane observed is formed by the following reaction.



(A) (B) (C) (D) (E)

How to Reach Us

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