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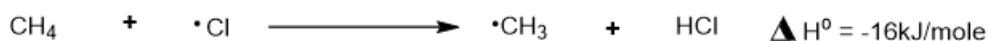
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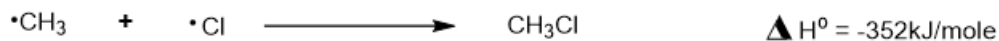
Choose the **incorrect** statement about the generally accepted free radical reaction mechanism for the chlorination of methane.



a) The change in enthalpy of the following propagation step is -16 kJ/mol.



b) The change in enthalpy for the following termination step is -352 kJ/mol.



- c) Chlorine atoms are produced in the initiation step.
d) Methyl radicals are produced in one of the propagation steps.
e) Only a small fraction of the chloromethane observed is formed by the following reaction.



☐ (A) ☐ (B) ☐ (C) ☐ (D) ☐ (E)

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